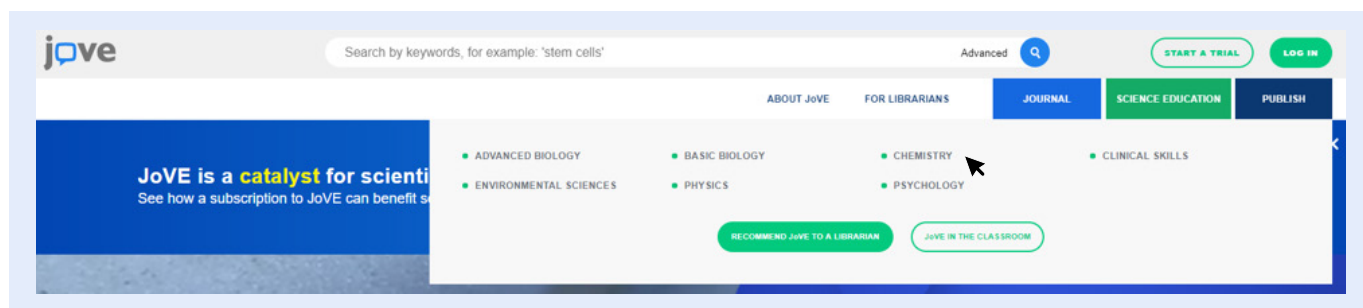
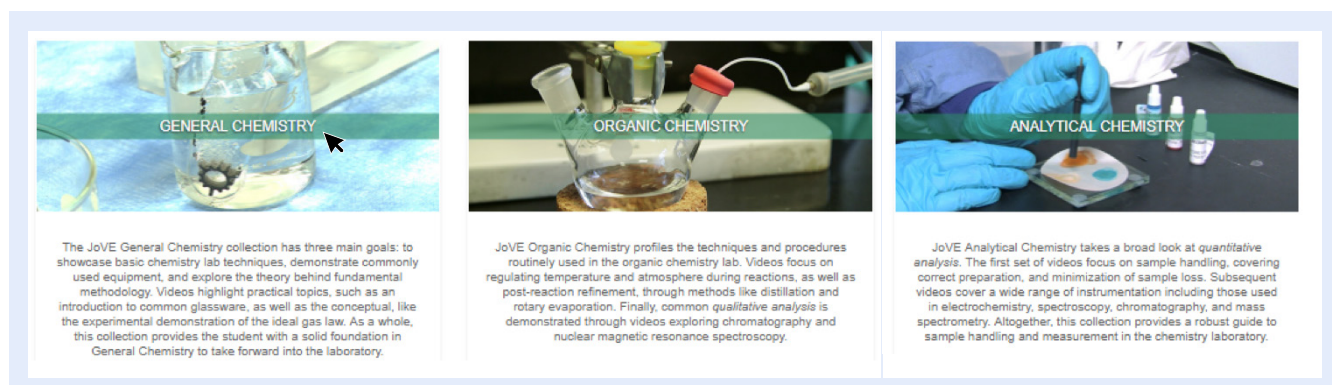


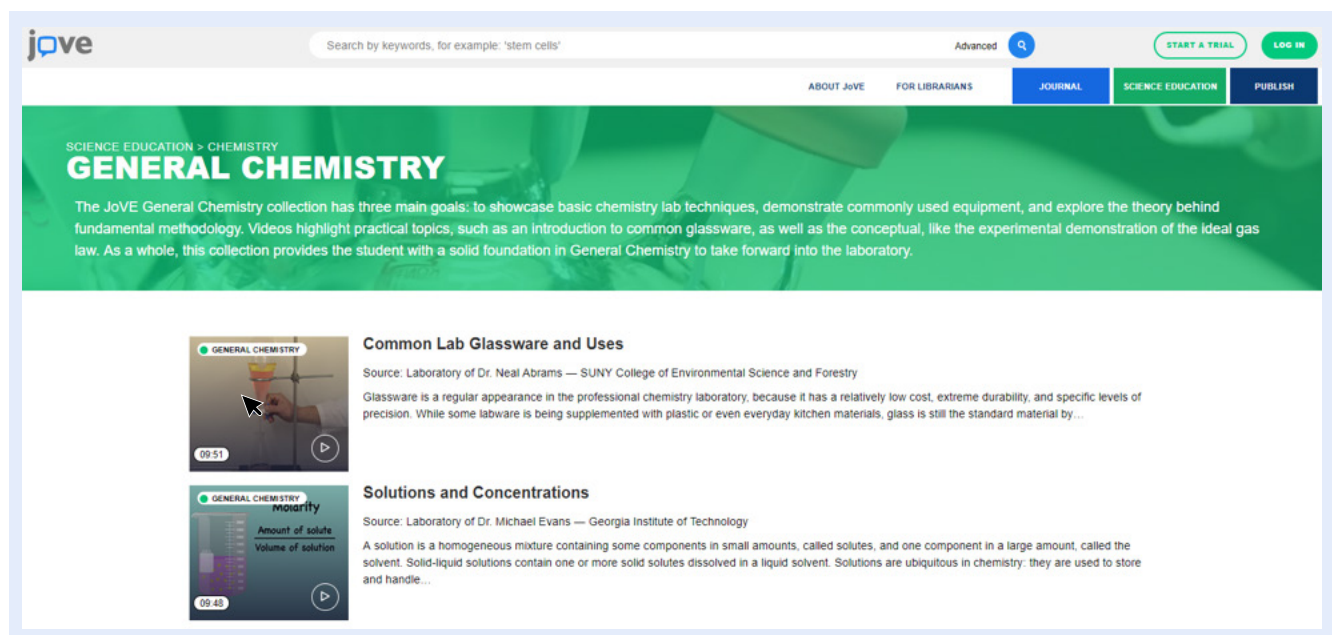
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- 0:00 Overview

Solutions are ubiquitous in chemistry. They are used to store and handle small amounts of material, carry out chemical reactions, and develop materials with controllable properties. A solution is a homogeneous mixture containing some components in small amounts, called solutes, and one component in a large amount, called the solvent. The amount of solute relative to the total amount of a solution is known as its "concentration". Depending on whether it is the mass, volume, or amount in moles of the solution components that are being considered, this measurement can be expressed in a number of different ways, according to the needs of the experiment. In this video, we will first review the different types of units for measuring a solution's concentration. We will then go through a protocol for
- 1:07 Principles of Solution Concentration
- 2:56 Making a Sucrose Solution
- 4:42 Making a Supersaturated Sucrose Solution
- 6:34 Applications
- 9:18 Summary

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